

**THE SEVENTH YUGOSLAV MATERIALS
RESEARCH SOCIETY CONFERENCE**

YUCOMAT 2005

**Programme
and
The Book of Abstracts**

**HERCEG NOVI,
September 12-16, 2005**

Organized by:
YUGOSLAV MATERIALS RESEARCH SOCIETY
and
INSTITUTE OF TECHNICAL SCIENCES OF SASA

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INTRODUCTION:

Materials science and engineering incorporate acquiring of knowledge on synthesis and processing of materials, their composition and structure, properties and behaviour, functions and potentialities as well as application of that knowledge to various final products. Economic prosperity, life quality, and healthy environment are tightly connected with the improvements in the existing and the development of new materials and processing technologies. These improvements and development can contribute greatly to the national priorities: energy saving, environment and health protection, information and communication, infrastructure, transportation, etc.

The Yugoslav Materials Research Society (Yu-MRS), a non-government and non-profit scientific association, was founded in 1997 to promote multidisciplinary goal-oriented research in materials science and engineering. Main task and objective of the Society is to encourage creativity in materials research and engineering to reach a harmonic coordination between achievements in this field in our country and analogous activities in the world with an aim to include our country into the global international projects.

The First Conference on materials science and engineering, including physics, physical chemistry, condensed matter chemistry, and technology in general, was held in September 1995, in Herceg Novi. An initiative to establish Yugoslav Materials Research Society was born at the conference. Similar to other MR societies in the world, the programme was made and objectives determined. Until 2003 Conferences were held every second year and then they grew into Annual Conference that will be traditionally held in Herceg Novi in September of every year.

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EQUILIBRIUM AND KINETIC STUDIES OF HEAVY METAL IMMOBILISATION BY HYDROXYAPATITE

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Sorption capacities and kinetics of toxic heavy metals sorption by synthetic, well-characterized, hydroxyapatite (HAP) was studied using the batch technique. Initial concentration of Pb^{2+} , Zn^{2+} , Cd^{2+} and Sr^{2+} solutions, obtained from their nitrate salts, was $2 \cdot 10^{-3}$ mol/dm³ for kinetic experiments. The effect of initial metal concentration (10^{-4} - 10^{-2} mol/dm³ for Zn^{2+} , Cd^{2+} and Sr^{2+} and 10^{-4} - $5 \cdot 10^{-2}$ mol/dm³ for Pb^{2+}) on the amounts of retained heavy metals was examined in order to calculate sorption capacities. Initial pH values, in all batches were adjusted to 5.0 ± 0.1 . Suspensions of HAP and individual metal solutions were equilibrated on a horizontal shaker, at the solid to solution ratio of 1:200, at room temperature ($20 \pm 1^\circ\text{C}$). The suspensions were filtrated and analyzed for residual heavy metal concentration, final pH and the amount of Ca^{2+} released into the solution at different time intervals (from 5 minutes up to 2 days) for kinetic and at 24 h for equilibrium studies. Sorption of all examined metals was followed by Ca^{2+} release from crystal lattice of HAP, indicating ion-exchange or dissolution-precipitation mechanisms. The isotherm data were found to be well fitted by linear Langmuir equation. The sorption of Pb^{2+} was found to be the most rapid in respect to other cations (for equilibrium, a contact time of approximately 30 minutes was required for Pb^{2+} and 24 hours for Cd^{2+} , Zn^{2+} , and Sr^{2+}). In order to determine the sorption mechanisms and rate constants, two most widely used kinetics models in sorption processes (pseudo-first and pseudo-second order equations) have been applied to experimental data. According to correlation coefficients and calculated values of equilibrium concentrations, the pseudo-second order model was found to be more appropriate, suggesting that the rate limiting step in heavy metal sorption on HAP involves valence forces through sharing or exchange of electrons between the sorbent and sorbates.